Extending the Kinetic and Thermodynamic Limits of Molecular-Beam Epitaxy Utilizing Suboxide Sources or Metal-Oxide-Catalyzed Epitaxy

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We observe a catalytic mechanism during the growth of III-O and IV-O materials by suboxide molecular-beam epitaxy (S-MBE). By supplying the molecular catalysts In₂O and SnO we increase the growth rates of Ga₂O₃ and In₂O₃. This catalytic action is explained by a metastable adlayer A, which increases the reaction probability of the reactants Ga₂O and In₂O with active atomic oxygen, leading to an increase of the growth rates of Ga₂O₃ and In₂O₃. We derive a model for the growth of binary III-O and IV-O materials by S-MBE and apply these findings to a generalized catalytic description for metal-oxide-catalyzed epitaxy (MOCATAXY), applicable to elemental and molecular catalysts. We introduce a mathematical description of S-MBE and MOCATAXY, providing a computational framework to set growth parameters in previously inaccessible kinetic and thermodynamic growth regimes when using the aforementioned catalysis. Our results indicate that MOCATAXY takes place with a suboxide catalyst rather than with an elemental catalyst. As a result of the growth regimes achieved, we demonstrate a Ga₂O₃/Al₂O₃ heterostructure with an unrivaled crystalline quality, paving the way for the preparation of oxide device structures with unprecedented perfection.

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I. INTRODUCTION

Molecular-beam epitaxy (MBE) takes place under nonequilibrium conditions and surface kinetics plays a dominant role in the MBE growth process—allowing growth modes to be intentionally manipulated [1–5]. For decades, the single-step reaction mechanism occurring during the MBE growth of III-V (e.g., GaAs, GaN, AlN) [6–8] and II-VI (e.g., ZnSe, ZnO) [9,10] compound semiconductors has defined MBE as a rather simple and straightforward thin-film technique, especially when compared with chemical-vapor-deposition methods [11]. In contrast to the growth of III-V and II-VI materials, the surface kinetics of III-O (e.g., Ga₂O₃ and In₂O₃) [12–18] and IV-O (e.g., SnO₂) [19] compounds is governed by a complex reaction pathway, resulting in a two-step reaction mechanism to form the intended compound. The formation and subsequent desorption of admolecules, called suboxides (e.g., Ga₂O, In₂O, SnO), define the growth-limiting step for these classes of materials. The result is a rather narrow growth window within the adsorption-controlled regime [14,15,18–20].

By engineering the MBE processes, two variants of MBE have recently been developed. These variants extend the kinetic and thermodynamic limits within which III-O and IV-O materials may be grown. The first variant, suboxide MBE (S-MBE) [21], refers to a technique that uses the decomposition of III-VI and IV-VI compounds (i.e., MeₓOₓ) by group-III and group-IV elements [13,17] as well as a special MBE source chemistry [21] to produce suboxide molecular beams that consist almost entirely (typically > 99.9%) of a single suboxide molecular species (i.e., MeₓOₓ−ₓ). Stoichiometric coefficients of MeₓOₓ and MeₙOₓ−ₓ are x = 2 and y = 3 for III-O (e.g., Ga₂O₃, In₂O₃) and x = 1 and y = 2 for IV-O (e.g., SnO₂) materials.

Using the S-MBE approach, the growth-limiting step occurring during the growth of III-O and IV-O materials by conventional MBE is bypassed, enabling the growth of films with excellent structural quality and surface smoothness at growth rates exceeding 1 µm h⁻¹.
and a comparatively low growth temperature \((T_G)\) [21]. The second variant is metal-oxide-catalyzed epitaxy (MOCATAXY), a method involving the introduction of a catalyst into oxide growth systems [5,17,18,22]. Using this technique, the growth-limiting step of III-O compounds is bypassed by the catalyst [5]. It has been proposed that MOCATAXY results from metal-exchange catalysis (MEXCAT) [5] and it has been observed on various growth surfaces for the formation of \(\text{Ga}_2\text{O}_3\) and \((\text{Al}_x\text{Ga}_{1-x})_2\text{O}_3\) [5,18,22–25] as well as during other physical vapor-deposition methods [26]. Nevertheless, the underlying physics leading to the observed catalysis has been explained differently—experimentally [5,22] and theoretically [27]—and thus remains disputable.

In this paper, the combination of \(S\)-MBE and MOCATAXY is investigated and shown to result in a marked extension of the kinetic and thermodynamic limits of the growth of \(\text{Ga}_2\text{O}_3\) and \(\text{In}_2\text{O}_3\). By supplying suboxide molecular beams of \(\text{Ga}_2\text{O}_x\), \(\text{In}_2\text{O}_x\), and \(\text{SnO}\), a growth-rate \((\Gamma)\) enhancement of \(\text{Ga}_2\text{O}_3\) and \(\text{In}_2\text{O}_3\) is observed. Quantitative models describing this enhanced growth rate during \(S\)-MBE as well as during the catalyzed \(S\)-MBE of III-O thin films are derived and applied to the experimental observations to extract model parameters. Through a systematic comparison of experimental \(\Gamma\) data from different growth systems and growth methods (i.e., MOCATAXY during \(S\)-MBE and conventional MBE), a generalization of the proposed MEXCAT mechanism [5] is developed. This generalized growth mechanism is applicable to the growth of \(\text{Ga}_2\text{O}_3\) and \(\text{In}_2\text{O}_3\) by conventional MBE and \(S\)-MBE.

**II. SUBOXIDE-MBE (\(S\)-MBE) MODEL**

We begin by deriving a growth-rate model for the growth of III-VI and IV-VI compounds with general formula \(\text{Me}_x\text{O}_y\) by \(S\)-MBE and validate it by using \(\text{Ga}_2\text{O}_3\) and \(\text{In}_2\text{O}_3\) as examples. Figure 1 shows the growth kinetics of \(\text{Ga}_2\text{O}_3\) and \(\text{In}_2\text{O}_3\) as a function of their respective growth parameters. In Figs. 1(a)–1(d), the film growth rate, \(\Gamma\), is observed to increase linearly with the incident flux of suboxide, \(\phi_{\text{Me}_x\text{O}_y}\) (i.e., \(\phi_{\text{Ga}_2\text{O}_3}\) or \(\phi_{\text{In}_2\text{O}_3}\)), in the O-rich regime. This linear increase in \(\Gamma\) reaches a plateau in the adsorption-controlled regime, when \(\phi_{\text{Me}_x\text{O}_y}\) exceeds the flux of active atomic oxygen, \(\phi_o\). The growth kinetics for the growth of III-O compounds by \(S\)-MBE are thus the same as those of III-V and II-VI materials when grown by conventional MBE [21].

Figures 1(e) and 1(f) depict \(\Gamma\) as a function of \(T_G\) and \(\phi_o\), respectively. For the same growth conditions, \(\Gamma\) of \(\text{Ga}_2\text{O}_3\) on \(\text{Ga}_2\text{O}_3(010)\) (hollow squares) is larger than that of \(\text{Ga}_2\text{O}_3\) on \(\text{Ga}_2\text{O}_3(201)\) (solid squares) [see Fig. 1(e)]. This result is similar to the growth of \(\text{Ga}_2\text{O}_3\) by conventional MBE on \(\text{Ga}_2\text{O}_3(201)\) versus on \(\text{Ga}_2\text{O}_3(010)\) substrates [20,28,29]. A comparison of the growth kinetics of \(\text{Ga}_2\text{O}_3(201)\) and \(\text{In}_2\text{O}_3(111)\) [e.g., the data in Fig. 1(a) with the data in Fig. 1(d)] establishes that the range of \(T_G\) within which high-quality films of \(\text{In}_2\text{O}_3\) can be grown at high \(\Gamma\) is larger than that for \(\text{Ga}_2\text{O}_3\). This result is also similar to the growth of \(\text{Ga}_2\text{O}_3\) and \(\text{In}_2\text{O}_3\) by conventional MBE [13,15,18].

To model the growth of binary oxides (\(\text{Me}_x\text{O}_y\)) from their suboxides (\(\text{Me}_x\text{O}_{y-\gamma}\)), we take the single-step
reaction kinetics of S-MBE [21] into account. Here, the growth takes place via the reaction
\[
\text{Me}_x\text{O}_{y-4}(a) + x\text{O}(a) \rightarrow \text{Me}_x\text{O}_{y}(s),
\]
with constant \( \kappa \) describing the \( \text{Me}_x\text{O}_y \) formation rate. Adsorbate and solid phases are denoted as \( a \) and \( s \), respectively. Based on the reaction given in Eq. (1), we set up a generalized growth-rate model describing the growth of these materials by S-MBE:
\[
\frac{dn_r}{dt} = \phi_r - \kappa n_r n_o^\delta - \tau_r^{-1} n_r, \quad (2)
\]
\[
\frac{dn_o}{dt} = \phi_o - x\kappa n_r n_o^\delta - \tau_o^{-1} n_o, \quad (3)
\]
\[
\frac{dn_p}{dt} = \Gamma = \kappa n_r n_o^\delta. \quad (4)
\]
The surface densities of the adsorbed cationic reactants (\( r = \text{Me}_x\text{O}_{y-4} \)), O adsorbates, and formed products (\( p = \text{Me}_x\text{O}_y \)) are denoted as \( n_r, n_o, \) and \( n_p \), respectively. Their time derivative is described by the operator \( d/dt \). The surface lifetimes of \( r \) and \( O \) are given by \( \tau_r \) and \( \tau_o \), respectively. The \( T_G \)-, orientation-, and growth-system-dependent sticking probability \( \zeta \) of O can be derived by considering two competing processes: the chemisorption of O and the desorption of O from the given adsorption sites, described by the reaction rates \( k_i \) and \( k_d \), respectively,
\[
k_i = v_i e^{-\frac{\delta \zeta}{k_B T}},
\]
with \( i = c, d \), frequency factor \( v_i \), and the respective energy barrier \( \epsilon_i \). The Boltzmann constant is denoted \( k_B \). Following the Kisliuk model [30], the O sticking probability \( 0 \leq \zeta \leq 1 \) [31,32] is given by
\[
\zeta = \frac{k_c}{k_c + k_d} = \left( 1 + ve^{-\frac{\delta \zeta}{k_B T}} \right)^{-1},
\]
with dimensionless prefactor \( v = v_d/v_c \) and activation barrier \( \delta \epsilon = \epsilon_d - \epsilon_c \). In the case of a high O desorption barrier or a high O adsorption barrier, the limits of Eq. (6) are (i) \( \epsilon_d \gg \epsilon_c \Rightarrow \delta \epsilon \gg 0 \Rightarrow \zeta \rightarrow 1 \) and (ii) \( \epsilon_c \gg \epsilon_d \Rightarrow \delta \epsilon \ll 0 \Rightarrow \zeta \rightarrow 0 \), respectively, satisfying our conditions of \( 0 \leq \zeta \leq 1 \). Figure 2 depicts the model results of Eq. (6) of our model for \( \text{Ga}_2\text{O}_3 \) and \( \text{In}_2\text{O}_3 \) using the parameters given in Table I.

Solving Eqs. (2) and (3) with respect to \( n_r \) and \( n_o \) and inserting their solutions into Eq. (4) yields an analytical expression for \( \Gamma \) described by three kinetic parameters: \( \kappa, \tau_r \), and \( \tau_o \) [see Eqs. (A1)–(A7)]. To reduce the complexity of the model, we assume that O adsorbate desorption is negligible and use \( \kappa \gg \tau_r^{-1} \) and \( \tau_o \ll \tau_r \). Assuming that the observed reaction processes follow a thermally activated Arrhenius behavior, we further reduce the complexity of our model by forming the product
\[
P(T_G) = \tau_c \kappa = P_0 e^{\left( \frac{\zeta}{k_B T} \right)},
\]
with preexponential factor \( P_0 = \tau_c^0 \kappa_0 \) and activation energy \( \epsilon = \epsilon_c + \epsilon_\kappa \). The latter expression yields a larger \( \kappa \) or a longer \( \tau_r \), the higher the growth rate \( \Gamma \) of the intended compound. This is in agreement with experiment as shown in Fig. 1.

We apply the solution of this model to the binary growth rate data of \( \text{Ga}_2\text{O}_3 \) and \( \text{In}_2\text{O}_3 \) by S-MBE depicted in Fig. 1 and extract the kinetic parameters summarized in Table I. To extract the parameters, we use an iterative approach. For example, for the growth of \( \text{Ga}_2\text{O}_3 \) by S-MBE, we first establish that the functional form of the equations accurately describes the growth of (010)-oriented \( \beta\)-\( \text{Ga}_2\text{O}_3 \) films, as plotted in Figs. 1(a) and 1(b). Having established that the functional form of the equations [solutions given by Eqs. (A1)–(A7)] accurately describes the growth of (010)-oriented \( \beta\)-\( \text{Ga}_2\text{O}_3 \) films, the model is next expanded to an additional orientation, (010) \( \beta\)-\( \text{Ga}_2\text{O}_3 \) films, using the data in Fig. 1(c).

A. Orientation-dependent growth rate of \( \beta\)-\( \text{Ga}_2\text{O}_3 \)
As a quantitative result, we find that the range of \( T_G \) that can be used to produce high-quality epitaxial films at

<table>
<thead>
<tr>
<th>( \beta)-( \text{Ga}_2\text{O}_3(201) )</th>
<th>( \beta)-( \text{Ga}_2\text{O}_3(010) )</th>
<th>( \text{In}_2\text{O}_3(111) )</th>
</tr>
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<tr>
<td>( \nu )</td>
<td>( \delta \epsilon ) (eV)</td>
<td>( P_0 ) (nm²s⁻¹)</td>
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<tr>
<td>( 1.07 \pm 0.1 )</td>
<td>( 0.15 \pm 0.05 )</td>
<td>( 0.12 \pm 0.01 )</td>
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<tr>
<td>( 1.20 \pm 0.2 )</td>
<td>( 0.25 \pm 0.09 )</td>
<td>( 0.15 \pm 0.03 )</td>
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<tr>
<td>( 3.21 \pm 0.5 )</td>
<td>( 0.90 \pm 0.10 )</td>
<td>( 0.25 \pm 0.02 )</td>
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FIG. 2. The O sticking probability \( \zeta \) as a function of the growth temperature \( T_G \). The solid, dashed, and dashed-dotted lines represent \( \zeta \) of \( \beta\)-\( \text{Ga}_2\text{O}_3(201) \) (the dark-blue area), \( \beta\)-\( \text{Ga}_2\text{O}_3(010) \) (the blue area), and bixbyite (cubic) \( \text{c-In}_2\text{O}_3(111) \) (the pale-blue area), respectively. The lines are modeled by Eq. (6) using the parameters given in Table I for the respective phases and surface orientations.
high Γ (i.e., a growth window) of (201)-oriented Ga$_2$O$_3$ films is narrower than for the growth of (010)-oriented Ga$_2$O$_3$ films, which, in turn, is narrower than the one for the growth of (111)-oriented In$_2$O$_3$ films. This is because of the different reaction efficiencies, η, of Ga$_2$O and In$_2$O with active O species and resulting O reservoirs, depending on the specific growth surface and the specific growth system (more detailed explanation below in the text). In addition, we explain the different sticking probabilities ζ, e.g., as obtained for Ga$_2$O$_3$(201) and Ga$_2$O$_3$(010), by the activity of surface reactions between Ga$_2$O and O adsorbates being dependent on the orientation of the surface on which the reaction takes place. In other words, the reservoir of active atomic oxygen for Ga$_2$O on which the reaction takes place. In other words, the reservoir of active atomic oxygen for Ga$_2$O on this surface orientation for the growth of Ga$_2$O by conventional MBE, where a similar dependence of Ga$_2$O growth rate is observed. We propose that this Γ dependence results from an orientation-dependent activity of the oxidation of Ga$_2$O to Ga$_2$O$_3$ [14,20] and that the underlying reason is due to the orientation-dependent vertical and lateral bond strengths between ad-atoms and the substrate surfaces [28,29]. In Refs. [20,28,29], the orientation dependence of the growth rate on the (hkl) plane, Γ(hkl), is given for β-Ga$_2$O$_3$: Γ(010) > Γ(001) > Γ(201) > Γ(100). We observe this same order of growth rate as a function of surface orientation for the growth of Ga$_2$O$_3$ by S-MBE and conclude that the orientation-dependent values of the sticking probability ζ(010) > ζ(001) > ζ(201) > ζ(100) can be related to the orientation-dependent reaction activities of adsorbates, underlying the observed order of Γ(hkl) for β-Ga$_2$O$_3$ (see Fig. 2).

III. MOCATAXY COMBINED WITH S-MBE

We next describe the enhancement of the growth rate of Ga$_2$O$_3$ and In$_2$O due to the presence of the catalysts In$_2$O and SnO, i.e., the combined effects of S-MBE and MOCATAXY. As depicted in Figs. 3(a)–3(c), we observe a drastic enhancement in Γ of heteroepitaxial β-Ga$_2$O$_3$(201) grown on Al$_2$O$_3$(0001) and homoepitaxial β-Ga$_2$O$_3$(010) grown on β-Ga$_2$O$_3$(010) using the suboxides In$_2$O and SnO as catalysts. The catalytic effect of SnO on the growth rate of Ga$_2$O$_3$ is stronger than that of In$_2$O. It is also stronger for the growth of β-Ga$_2$O$_3$(201) (solid symbols) than for the growth of β-Ga$_2$O$_3$(010) (hollow symbols). The stronger catalytic effect of SnO compared with In$_2$O can be explained by their different vapor pressures, i.e., $P_{\text{SnO}} < P_{\text{In}_2\text{O}}$ [33–35]. Thus, under similar growth conditions, the surface lifetime of SnO is longer than that of In$_2$O. Hence, SnO can be reoxidized more often than In$_2$O [5,17] and the O reservoir that ultimately ends up oxidizing Ga$_2$O in the presence of SnO is larger than that for Ga$_2$O oxidation in the presence of In$_2$O. Figures 3(d)–3(f) depict Γ of Ga$_2$O$_3$ as a function of $\phi_{\text{In}_2\text{O}}$ and $\phi_{\text{SnO}}$, respectively. When SnO is used as a catalyst, it is evident from the data that the catalytic effect on Ga$_2$O$_3$(201) is stronger than it is on Ga$_2$O$_3$(010). As shown below in Eq. (16), the catalytic activity, $\alpha$, decreases with an increasing O flux. Therefore, the weaker catalytic effect observed on β-Ga$_2$O$_3$(010) compared with the effect on β-Ga$_2$O$_3$(201)
may be explained by the higher surface-dependent sticking probability $s_{(010)} > s_{(201)}$ (values are given in Table I). Further studies on the atomic incorporation and surface segregation of residual Sn and In in the grown thin films, when using SnO and In$_2$O as catalysts, respectively, need to be performed. Nevertheless, energy-dispersive x-ray spectroscopy (EDXS), secondary-ion mass spectrometry (SIMS), and atomic probe tomography (APT) have already shown that the concentrations of In [5,18,24] and Sn [22] in Ga$_2$O$_3$ as well as in (Al$_x$Ga$_{1-x}$)$_2$O$_3$ films are below 1%. This is consistent with our x-ray diffraction (XRD) results, showing only diffraction peaks from Ga$_2$O$_3$ in the thin films, an example of which is shown in Fig. 4.

A comparison of Figs. 3(f) and 3(g) reveals that the growth rates $\Gamma$ of $\beta$-Ga$_2$O$_3$(201) and $\beta$-Ga$_2$O$_3$(010) using solely SnO and SnO + In$_2$O as catalysts are very similar (under otherwise identical growth conditions). We thus conclude that SnO suppresses In$_2$O as a catalyst, i.e., the catalytic effects are not additive: the presence of SnO inhibits (or isolates) the catalytic activity of In$_2$O in the SnO—In$_2$O—Ga$_2$O$_3$ system. The prediction of this same effect for MOCATAXY by conventional MBE, i.e., when using the elemental catalysts In [5] and Sn [22] in combination for the growth of Ga$_2$O$_3$ in the Sn—In—Ga—O system.

Figure 3(h) shows the catalytic effect of SnO on $\Gamma$ of In$_2$O$_3$(111) grown on Al$_2$O$_3$(0001) substrates. The solid circle in Fig. 3(h) corresponds to the growth rate of In$_2$O$_3$ determined by x-ray reflectivity (XRR) in the absence of SnO. In the presence of SnO, the growth rate of In$_2$O$_3$ (again measured by XRR) is seen to increase by a factor of approximately 1.4, i.e., far larger than the effect caused by additional SnO incorporation into In$_2$O$_3$. SIMS [37] and x-ray fluorescence (XRF) [37] are used to quantify the Sn content present in the In$_2$O$_3$ films grown with the SnO catalyst. A value of approximately 3% Sn in In$_2$O$_3$ is observed by SIMS and XRF independent of the SnO fluxes in the (0.2–0.5) nm$^{-2}$ s$^{-1}$ range.

This result indicates that SnO increases the available O reservoir for the oxidation of In$_2$O$_3$ to In$_2$O$_3$ by a factor of approximately 1.4 (at these growth conditions). This finding is in line with the data plotted in Figs. 3(f) and 3(g), i.e., that SnO has a higher reactivity with O than In$_2$O. We explain the catalytic action of SnO on In$_2$O by their different vapor pressures $P_{\text{SnO}} < P_{\text{In}_2\text{O}}$ [33–35]. Thus, SnO provides a larger O reservoir for In$_2$O oxidation than is available in the absence of SnO. The In$_2$O—SnO—O system is the first catalytic system observed for MBE growth beyond Ga$_2$O$_3$-based systems [5,18,22–25]. The discovery of MOCATAXY during S-MBE and its extension to In$_2$O$_3$ suggests its universality for the MBE growth of a multitude of oxide compounds (during S-MBE as well as conventional MBE).

We explain the observed catalysis during the growth of Ga$_2$O$_3$ and In$_2$O$_3$ by S-MBE by the formation of a metastable adlayer $A$ between the catalyst $c$ and O, e.g., $A = \text{In}_2\text{O} – \text{O} = \text{SnO} – \text{O}$. As experimentally and mathematically shown by Fig. 1, Table I, and Eqs. (2)–(6), In$_2$O possesses a higher surface reactivity with active O than does Ga$_2$O$_3$, leading to a higher growth rate of In$_2$O$_3$ compared with Ga$_2$O$_3$ at comparable growth conditions. This behavior is similar to the growth of Ga$_2$O$_3$ and In$_2$O$_3$ by conventional MBE [13,16], which has been explained by the different oxidation efficiencies $\eta$ of the elements Sn, In, and Ga following the order $\eta_{\text{Sn}} \approx 1.1 \eta_{\text{In}} \approx 3.1 \eta_{\text{Ga}}$ [5,13,17]. Taking the ratio of maximum available O for Ga$_2$O$_3$ and In$_2$O$_3$ oxidation [data plotted in Figs. 1(b) and 1(d) and flux conversions given in Eqs. (A11)] from mixtures of O$_2$ and 10% O$_3$ [38] at a background pressure of $1 \times 10^{-6}$ Torr, we obtain $\eta_{\text{In}_2\text{O}_3} \approx 2.8 \eta_{\text{Ga}_2\text{O}_3}$. This result is very similar to $\eta_{\text{In}} \approx 2.8 \eta_{\text{Ga}}$ [5,13,17] observed during conventional MBE growth. We surmise, following Ref. [17], that the same value of $\eta$ observed during MOCATAXY in conventional MBE and S-MBE arises from what is the second reaction step of conventional MBE and the sole reaction step of S-MBE, where the suboxide reacts with O to complete the formation of the intended oxide (e.g., Ga$_2$O$_3$ → Ga$_2$O$_3$) [18] and not to the suboxide-formation step (e.g., Ga → Ga$_2$O) [18]. Thus, for the discussion and analysis that follow, we use $\eta_{\text{SnO}} \approx 1.1 \eta_{\text{In}_2\text{O}_3} \approx 3.1 \eta_{\text{Ga}_2\text{O}_3}$ during S-MBE. These oxidation efficiencies follow the same order in $\eta$ as has been observed for growth by conventional MBE [5,17].

A. Generalized metal-exchange catalytic model

We propose that the role of the catalyst $c$ (e.g., In$_2$O, SnO, In [5], or Sn [22]) is to increase the O adsorbate reservoir of the reactant $r$ (e.g., Ga$_2$O$_3$ or Ga) by forming $A$ through the reaction

$$c(a) + O(a) \rightarrow A(a), \quad (8)$$

with examples of $A$ being In$_2$O—O, In—O, SnO—O, or Sn—O. In the presence of $r$, $A$ is unstable and catalyzes the incorporation of $r$ into the intended product $p$ (e.g., Ga$_2$O$_3$), while decreasing the reaction barrier of $r$ with O. Thus, the reaction given in Eq. (8) is subsequently followed by the reaction

$$\text{Me}_x\text{O}_{y-x}(a) + xA(a) \rightarrow \text{Me}_x\text{O}_y(s) + c(a). \quad (9)$$

Equation (9) describes the consumption of $A$ while forming the product ($p = \text{Me}_x\text{O}_y$) and releasing $c$ on the growth surface. The catalyst $c$ may be reoxidized [5], leading to an increase in the available O reservoir for the reactant ($r = \text{Me}_x\text{O}_{y-x}$) and thus to an extension of the kinetic and thermodynamic limits to the formation of $p$.

An adlayer formed by In has also been observed during the formation of GaN using In as a surface active agent (surfactant) [3]. Here, the In adlayer enables an enhanced
difficulty channel for the Ga and N adsorbates. Moreover, a surface instability of In—O bonds in the presence of Fe [39] and Ga [5] has been observed during conventional MBE. We emphasize, however, that the catalytic effect we are describing must not be confused with effects resulting from surfactants during the growth of III-V compounds by conventional MBE [2−4].

In order to describe MOCATAXY for elemental (e.g., In [5] and Sn [22]) and molecular catalysts (e.g., In2O and SnO) as well as for different materials (e.g., Ga2O3 and In2O3) mathematically, we would have to take into account the surface populations of c, r, and atomic O together with the surface density of the p that forms following the Langmuir-Hinshelwood mechanism [40]. We may reduce the complexity of the model significantly by only taking into account the most likely reactions involved in the formation of p, following the Eley-Rideal formalism [41]. The resulting set of coupled differential equations reads as follows:

$$\frac{dn_c}{dt} = \gamma_c n_c (1 - \theta_A) \phi_o + \alpha \phi_r \theta_A - \gamma_c n_c, \quad (10)$$

$$\frac{d\theta_A}{dt} = \sigma n_c (1 - \theta_A) \phi_o - \alpha \phi_r \theta_A, \quad (11)$$

$$\frac{dn_o}{dt} = \Gamma = \alpha \phi_r \theta_A, \quad (12)$$

with the adatom density $n_c$ and desorption-rate constant $\gamma_c$ of the catalyst c, as well as the surface coverage $\theta_A$ of A. The second and third terms in Eq. (10) refer to the formation rates of $A$ and $p$, respectively, and the factor $(1 - \theta_A)$ assures that $A$ constitutes a surface phase. The last term in Eq. (10) accounts for the desorption of $c$ from the growth surface and $\sigma$ (with the dimension of nm$^2$) represents the cross section of colliding $c$ with $O$ [5]. The impinging fluxes of $c$ and $r$ are denoted as $\phi_c$ and $\phi_r$, respectively. We note that the structure of the model introduced here is similar to the MEXCAT model given in Ref. [5]. The improvement of the model given in this work and its generalization to elemental and molecular catalysts arises by taking a cationiclike, metastable adlayer $A$ into account. This allows MOCATAXY to be described for the growth of ternary systems involving molecular catalysts [e.g., In2O (Fig. 3) and SnO (Fig. 3)] as well as elemental catalysts [e.g., In (Fig. 5) and Sn (Fig. 5)].

### B. Suboxide catalysts

Our results thus indicate that MOCATAXY takes place with a suboxide catalyst (e.g., for $A$ being In$_2$O—O or SnO—O) and not with an elemental catalyst (e.g., for $A$ being In—O or Sn—O). For conventional MBE, i.e., when using elemental source materials, we assume that the reaction of the metal to form the suboxide (e.g., $2Ga + O \rightarrow Ga_2O = r$) occurs very rapidly [36] and, thus, the catalysis takes place between the suboxide reactant $r$ and $A$, satisfying the reactions given in Eqs. (8) and (9). Should the reaction of $2Ga + O \rightarrow Ga_2O$ not occur very rapidly, we would see Ga desorption (at least a fraction of it), and a plateau in $\Gamma$ in the adsorption-controlled regime during the growth of Ga$_2$O$_3$ by conventional MBE; similar to the $\Gamma$ plateau observed, e.g., during the growth of GaN by conventional MBE [8]. The subsequent desorption of the rapidly formed Ga$_2$O is the growth-rate-limiting step in conventional MBE, as it removes active O from the growth front (by forming Ga$_2$O), leading to the decrease in $\Gamma$ in the adsorption-controlled regime as well as at elevated $T_G$ [12,17,36]. In S-MBE, the formation of Ga$_2$O (through $2Ga + O \rightarrow Ga_2O$) is bypassed since Ga$_2$O is directly provided from the source. Therefore, the O-consuming step is now avoided and a plateau in $\Gamma$ of Ga$_2$O$_3$ in the adsorption-controlled regime occurs [21]. The assumption that MOCATAXY takes place with a suboxide catalyst is further supported by our finding in this work that the oxidation efficiency $\eta$ for In$_2$O and Ga$_2$O during MOCATAXY by S-MBE follows the same order in $\eta$ as for Ga and In during conventional MBE, i.e., $\eta_{In_2O} = 2.8\eta_{Ga_2O}$ for S-MBE (this work) versus $\eta_{In} = 2.8\eta_{Ga}$ for conventional MBE [5,17]. Nevertheless, we emphasize that the microscopic origin and reaction pathways of MOCATAXY during S-MBE and conventional MBE require further investigation to fully understand which species are indeed involved, leading to the observed catalysis given in this work and presented in Refs. [5,17,18,22−27].

The solution of Eqs. (10) and (11) with respect to $n_c$ and $\theta_A$ yields $\phi_c = \gamma c n_c$, consistent with our observation of negligible incorporation (approximately 3%) of $c$ into the grown thin films for the data plotted in Fig. 3. Insertion of the solution for $\theta_A$ into Eq. (12) yields the following expression:

$$\Gamma = \frac{\alpha \phi_r \phi_o}{\alpha J \phi_r + \phi_o \phi_r}, \quad (13)$$

The free parameters are the pseudoflux

$$J = \gamma c = J_0 \exp \left( -\frac{\Delta}{k_BT_G} \right), \quad (14)$$

with $J_0 = 1 \times 10^{14}$ nm$^{-2}$ s$^{-1}$ (assumed for all species as a first approximation [31]) and energy $\Delta$. The value of $\Delta$ depends linearly on $\phi_c$ (e.g., $\phi_{In_2O}$ and $\phi_{SnO}$),

$$\Delta(\phi_c) = \Delta_0 + \delta \phi_c, \quad (15)$$

with $\Delta_0$ denoting the evaporation enthalpy, e.g., of In$_2$O and SnO, and $\delta$ describing its increase with increasing $\phi_c$. The other free parameter is the catalytic activity coefficient, $\alpha$. For $\alpha$ (ranging from 1 to 0), we use a linear
approximation depending on impinging $\phi_o$, i.e.,

$$\alpha(\phi_o) = 1 - b\phi_o, \quad (16)$$

with $b$ describing the decrease of $\alpha$ with $\phi_o$. By an iterative approach of our model to the experimental growth-rate data, the values obtained of $\Delta_0$, $\delta$, and $b$ for In$_2$O and SnO are given in Table II.

The evaporation enthalpies $\Delta_0$ of In$_2$O and SnO on Ga$_2$O$_3$ surfaces correspond to the values given in the literature, as noted in Table II. Only the value of $\Delta_0$ for SnO on the In$_2$O$_3$ surface is slightly above the literature value, indicating an additional energy term caused by vertical and lateral interactions between adsorbed species and the In$_2$O$_3$ surface. The increase in $\Delta$ with $\phi_o$, described by the parameter $\delta$, indicates an increase in the lateral adsorbate binding energy with increasing $\phi_o \propto n_o$.

Analogous behavior has also been observed for the desorption kinetics of In on Ga$_2$O$_3$ [5] and Ga on GaN [42]. We explain the linear decrease of $\alpha$ with $\phi_o$ by a linear increase of O adsorbates with $\phi_o$, i.e., $\phi_o \propto n_o$. It has been shown that an increase in $\phi_o$ promotes the incorporation of In during the growth of $(\text{In}_2\text{Ga}_{1-x})_2\text{O}_3$ by conventional MBE; this decreases the catalytic strength of In for the formation of Ga$_2$O$_3$ [5,16,17,29]. We therefore explain the decrease in $\alpha$ with increasing $\phi_o \propto n_o$ by a decreasing diffusion length of $c$ and $r$ and, thus, a decrease in the probability of the reaction given in Eq. (9) occurring.

### TABLE II. The values of $\Delta_0$, $\delta$, and $b$ used in Eqs. (15) and (16) for different catalytic systems and growth surfaces.

<table>
<thead>
<tr>
<th>System</th>
<th>$\Delta_0$ (eV)</th>
<th>$\delta$ (meV nm$^2$ s)</th>
<th>$b$ (nm$^2$ s)</th>
</tr>
</thead>
<tbody>
<tr>
<td>$\beta$-Ga$_2$O$_3$(201):In$_2$O</td>
<td>2.6 ± 0.05 [34]</td>
<td>30 ± 2</td>
<td>0.07 ± 0.01</td>
</tr>
<tr>
<td>$\beta$-Ga$_2$O$_3$(010):In$_2$O</td>
<td>2.6 ± 0.05 [34]</td>
<td>30 ± 2</td>
<td>0.07 ± 0.01</td>
</tr>
<tr>
<td>$\beta$-Ga$_2$O$_3$(201):SnO</td>
<td>2.9 ± 0.05 [33]</td>
<td>50 ± 4</td>
<td>0.20 ± 0.03</td>
</tr>
<tr>
<td>$\beta$-Ga$_2$O$_3$(010):SnO</td>
<td>2.9 ± 0.05 [33]</td>
<td>10 ± 1</td>
<td>0.33 ± 0.05</td>
</tr>
<tr>
<td>In$_2$O$_3$(111):SnO</td>
<td>3.2 ± 0.05</td>
<td>400 ± 10</td>
<td>0.33 ± 0.05</td>
</tr>
</tbody>
</table>

### D. Obtained “pseudo” $\beta$-Ga$_2$O$_3$/Al$_2$O$_3$

As a result of combining S-MBE with MOCATAXY, in Figs. 4(a)–4(d), we demonstrate an previously unknown “pseudo” $\beta$-Ga$_2$O$_3$(201)/Al$_2$O$_3$(0001) heterostructure with unparalleled crystalline perfection. Figure 4(a) shows $\theta$-2$\theta$ XRD scans of a 90-nm-thick SnO-catalyzed Ga$_2$O$_3$ film (black trace) and a 40-nm-thick noncatalyzed Ga$_2$O$_3$ film grown by S-MBE under similar growth conditions ($T_G = 500$ °C, $\phi_{Ga_3O} = 0.9$ nm$^{-2}$ s$^{-1}$, and $\phi_o = 2.3$ nm$^{-2}$ s$^{-1}$). The SnO-catalyzed Ga$_2$O$_3$ film is the same film that is plotted in Fig. 3(f) as a solid-black star. It is grown at $T_G = 690$ °C, $\phi_{Ga_3O} = 1.3$ nm$^{-2}$ s$^{-1}$, and $\phi_o = 2.3$ nm$^{-2}$ s$^{-1}$. The reflections of the noncatalyzed Ga$_2$O$_3$ thin film coincide with the $\beta$-Ga$_2$O$_3$(201) phase grown with its (201) plane parallel to the (0001) plane of the Al$_2$O$_3$ substrate. In contrast, the XRD scan of the SnO-catalyzed Ga$_2$O$_3$ film shows only the even reflections of the $\beta$-Ga$_2$O$_3$ phase grown with its (201) plane parallel to the (0001) plane of the Al$_2$O$_3$ substrate. We speculate that the “pseudo” $\beta$-Ga$_2$O$_3$(201)/Al$_2$O$_3$(0001) contains an periodic occurrence of low-energy stacking faults, parallel to the (201) plane of the Ga$_2$O$_3$ film, that reduce the number of the observed XRD diffraction peaks of the “pseudo” $\beta$-Ga$_2$O$_3$(201) phase. This hypothesis is consistent with calculations showing that the stacking fault energy in $\beta$-Ga$_2$O$_3$ is low [43,44], experimental observation of high densities of stacking faults in $\beta$-Ga$_2$O$_3$ films [45], and the scanning transmission electron microscopy (STEM) image depicted in Fig. 4(b). The epilayer shows a single-crystalline structure with an abrupt interface to the Al$_2$O$_3$(0001) substrate at the bottom of the image. Figure 4(c) shows a transverse scan (rocking curve) across the symmetric 402 reflection of the same “pseudo” $\beta$-Ga$_2$O$_3$(201) film. The measured full width at half maximum (FWHM) is just $\Delta_\omega \approx 10$ arcsec; this is a measure of the out-of-plane mosaic spread of the thin film. Figure 4(d) depicts an atomic force microscope (AFM) image of the surface morphology of the same SnO-catalyzed Ga$_2$O$_3$ film; it has a root-mean-square (rms) roughness of 2.5 Å. The obtained rocking curve of $\Delta_\omega \approx 10$ arcsec and smooth surface morphology of $\text{rms} = 2.5$ Å provide the best results obtained for any Ga$_2$O$_3$ thin film grown on Al$_2$O$_3$(0001) by any method [14,46–49]. We point out that optimizations utilizing the unprecedented growth regimes—becoming accessible by the combination...
of S-MBE with MOCATAXY—still need to be performed to further improve the crystalline perfection of the grown Ga$_2$O$_3$-based heterostructures.

We note that in the absence of a catalyst, Ga$_2$O$_3$ does not form at this high $T_G$ of 690 °C. In addition, at a lower catalyst flux of SnO [depicted as the open star in Fig. 3(f)] compared with the sample grown at a higher SnO flux [depicted as the solid star in Figs. 3(f) and 4], we do not measure the “pseudo” $\beta$-Ga$_2$O$_3$(201) phase, but instead see the “conventional” $\beta$-Ga$_2$O$_3$(201) peaks by XRD (data not shown in this work). We therefore conclude that extending the kinetic (e.g., higher possible $T_G$) and thermodynamic limits (e.g., different surface chemical potential) by combining S-MBE with MOCATAXY benefits the formation of the “pseudo” $\beta$-Ga$_2$O$_3$(201)/Al$_2$O$_3$(0001) heterostructure that has unparalleled crystalline perfection.

IV. CONCLUSION

As we demonstrate, the nature of the model derived to describe MOCATAXY using elemental and molecular catalysts does not depend on the specific growth surface. We note, however, that the growth surface may change the kinetic parameters used in our model, as shown for the examples of $\beta$-Ga$_2$O$_3$(201) and $\beta$-Ga$_2$O$_3$(010) using In$_2$O or SnO as catalysts in this work.

Finally, the increase in the growth rates of Ga$_2$O$_3$ and In$_2$O$_3$ by S-MBE that occurs when using the catalysts SnO and In$_2$O demonstrates MOCATAXY as a potentially inherent feature in conventional MBE growth [5,18,22–25] as well as in S-MBE growth. Furthermore, our results provide deeper insight into this catalysis, indicating that MOCATAXY occurs through a suboxide catalyst rather than with an elemental catalyst. This more broad applicability of MOCATAXY opens an unprecedented path for the epitaxial synthesis of thin films by intentionally extending the kinetic and thermodynamic limits during their growth processes. If successful, this could enable the growth of (yet) unknown crystal phases and unprecedented functional electronic materials.

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The values of \( \Delta_0 \) and \( \delta \) used in Eq. (14) (main text) for the catalytic model when using In and Sn as catalysts.

<table>
<thead>
<tr>
<th>Catalyst</th>
<th>( \Delta_0 ) (eV)</th>
<th>( \delta ) (meV nm(^2) s(^{-1}))</th>
</tr>
</thead>
<tbody>
<tr>
<td>Ga(_2)O(_3):In</td>
<td>2.42 ± 0.05 [50]</td>
<td>41 ± 2</td>
</tr>
<tr>
<td>Ga(_2)O(_3):Sn</td>
<td>2.91 ± 0.07 [33]</td>
<td>10 ± 2</td>
</tr>
</tbody>
</table>

Program (Grant No. DMR-1719875). Substrate preparation was performed, in part, at the Cornell NanoScale Facility, a member of the National Nanotechnology Coordinated Infrastructure (NNCI), which is supported by the NSF (Grant No. NNC1-2025233).

**APPENDIX**

We apply the generalized MOCATAXY model, Eqs. (10)–(16), to published catalytic data using elemental In [5] and elemental Sn [22] as catalysts for the growth of Ga\(_2\)O\(_3\) by MBE. The parameters used in Eqs. (15) for In and Sn are collected in Table III. Within experimental uncertainty, the obtained \( \Delta_0 \) corresponds to the vaporization enthalpies of elemental In [50] and Sn [33] as given in the literature; the values are given in Table III. Moreover, for our model, we use \( \alpha = 1 \) (the catalytic activity), the same value of \( \alpha \) as used in Ref. [5].

Figures 5(a) and 5(b) show our results for the growth of Ga\(_2\)O\(_3\) by MBE when supplying In and Sn as elemental catalysts, respectively. Our model drawn in Fig. 5(a) (solid lines) describes the data published in Ref. [5] with the same accuracy as the previously established MEXCAT model [5].

Figure 5(b) plots the catalyzed growth rate data of Ga\(_2\)O\(_3\) taken from Ref. [22]. Our model applied to this data is drawn as the solid line. It precisely describes the Sn-catalyzed growth of Ga\(_2\)O\(_3\) by MBE and it provides a quantitative description of MOCATAXY using Sn as a catalyst. To model the Sn-catalyzed data, we linearly extrapolate the value of \( \phi_0^S \) given in Eq. (A13) while using the O flux conditions published in Ref. [22].

To conclude, our model introduced in the main text is able to describe the growth of Ga\(_2\)O\(_3\) and In\(_2\)O\(_3\) by S-MBE as well as conventional MBE when using elemental and molecular catalysts. This is achieved by introducing a cationiclike, catalytic adlayer \( A \).

1. **Detailed model of the growth of III-VI materials by S-MBE**

The analytical solution of the growth rate, \( \Gamma \), of Eqs. (2)–(4) for III-VI compound materials is as follows:

\[
\Gamma = \left[ \xi_+^2 P^2 + \xi_-^2 - P \left( 3 - 2 \xi_+ \xi_- \right) \right] \times \left[ \xi_+^2 P^2 + \xi_-^2 - P \left( 3 - \xi_+ \xi_- \right) \right] \times \left( -54 P \xi_+ \right)^{-1},
\]

with

\[
\xi_+ = 2 \phi_r - \xi \phi_o, \tag{A2}
\]

\[
\xi_- = -2 \phi_r + \xi \phi_o, \tag{A3}
\]

\[
\xi_+ = \left( \xi_+ P - 9 \psi \psi - \sqrt{27 \psi} \right)^3, \tag{A4}
\]

\[
\xi_- = -\left( \xi_+ P + 9 \psi \psi + \sqrt{27 \psi} \right)^3, \tag{A5}
\]

\[
\psi = \phi_r + \xi \phi_o, \tag{A6}
\]

and

\[
\Psi = P^3 \left[ 1 + \xi_+^2 \xi_- \phi_0 + P \right. \times \left. \left( -\phi_r^3 + 10 \xi \phi_0 \phi_o + 2 \xi^2 \phi_o \right) \right]. \tag{A7}
\]

The growth rates of Ga\(_2\)O\(_3\) and In\(_2\)O\(_3\) as presented in Fig. 1 are explicitly modeled with Eqs. (A1)–(A7).

2. **Detailed model of the growth IV-VI materials by S-MBE**

The analytical solution for \( \Gamma \) of Eqs. (2)–(4) for IV-VI compound materials is as follows:

\[
\Gamma = \frac{1 + \sqrt{1 + 2 P \psi \psi + \psi^2 \psi^2}}{2P}, \tag{A8}
\]

with

\[
\psi = \phi_r \pm \xi \phi_o. \tag{A9}
\]

We predict that the \( \Gamma \) of IV-VI materials (e.g., SnO\(_2\))—obtained by S-MBE—may be modeled by Eqs.
(A8) and (A9) with the same accuracy as the \( \Gamma \) modeled for III-VI materials (e.g., Ga\(_2\)O\(_3\) and In\(_2\)O\(_3\)) by Eqs. (A1)–(A7); this is demonstrated in the main text.

3. Conversion factors for \( \phi_r \) and \( \phi_o \)

The model uses reactant fluxes, \( \phi_r \) (e.g., Ga\(_2\)O, In\(_2\)O, SnO, Ga, In, Sn) and active atomic oxygen fluxes, \( \phi_o \) (from active O\(_3\) or O species), in nm\(^2\) s\(^{-1}\).

The conversion factors for \( S \)-MBE used in an ozone MBE system are as follows:

\[
\begin{align*}
\phi_{\text{Ga}_2\text{O}}^\text{QCM} &= 1 \text{ Å s}^{-1} \rightarrow \phi_{\text{Ga}_2\text{O}} = 0.38 \text{ nm}^2 \text{s}^{-1}, \\
\phi_{\text{In}_2\text{O}}^\text{QCM} &= 1 \text{ Å s}^{-1} \rightarrow \phi_{\text{In}_2\text{O}} = 0.25 \text{ nm}^2 \text{s}^{-1}, \\
& \quad \text{(A10)} \\
\phi_{\text{SnO}}^\text{QCM} &= 1 \text{ Å s}^{-1} \rightarrow \phi_{\text{SnO}} = 0.45 \text{ nm}^2 \text{s}^{-1}, \\
P_{\text{Ga}_2\text{O}}^\text{O} &= 1 \times 10^{-6} \text{Torr} \rightarrow \phi_{\text{Ga}_2\text{O}} = 0.73 \text{ nm}^2 \text{s}^{-1}, \\
P_{\text{In}_2\text{O}}^\text{O} &= 1 \times 10^{-6} \text{Torr} \rightarrow \phi_{\text{In}_2\text{O}} = 2.05 \text{ nm}^2 \text{s}^{-1}, \\
P_{\text{SnO}}^\text{O} &= 1 \times 10^{-6} \text{Torr} \rightarrow \phi_{\text{SnO}} = 2.25 \text{ nm}^2 \text{s}^{-1}. \\
& \quad \text{(A11)}
\end{align*}
\]

The reactant fluxes, \( \phi_r^\text{QCM} \), are measured prior to growth by a quartz-crystal microbalance (QCM) in Å s\(^{-1}\), with the density of the QCM set to 1 g cm\(^{-3}\). These QCM readings are readily converted to absolute fluxes using the known masses of the Ga\(_2\)O, In\(_2\)O, and SnO molecules that condense onto the QCM. The active atomic oxygen fluxes given result from an oxidant, \( P_{\text{O}} \), with mixtures of O\(_2\) and approximately 10\% O\(_3\) in the oxygen molecular beam [38]. The different oxidation efficiencies are taken from Ref. [17] while using the results obtained for \( S \)-MBE.

The conversion factors for conventional MBE used in an oxygen-plasma-assisted MBE system are as follows:

\[
\begin{align*}
P_{\text{Ga}}^\text{BEP} &= 1 \times 10^{-7} \text{Torr} \rightarrow \phi_{\text{Ga}} = 1.2 \text{ nm}^2 \text{s}^{-1}, \\
P_{\text{In}}^\text{BEP} &= 1 \times 10^{-7} \text{Torr} \rightarrow \phi_{\text{In}} = 0.6 \text{ nm}^2 \text{s}^{-1}, \\
P_{\text{Sn}}^\text{BEP} &= 1 \times 10^{-7} \text{Torr} \rightarrow \phi_{\text{Sn}} = 1.5 \text{ nm}^2 \text{s}^{-1}, \\
\phi_{\text{O}}^\text{1 SCCM} &= 1 \text{ SCCM} \rightarrow \phi_{\text{O}}^\text{Ga} = 9.8 \text{ nm}^2 \text{s}^{-1}, \\
\phi_{\text{O}}^\text{1 SCCM} &= 1 \text{ SCCM} \rightarrow \phi_{\text{O}}^\text{In} = 27.4 \text{ nm}^2 \text{s}^{-1}, \\
\phi_{\text{O}}^\text{1 SCCM} &= 1 \text{ SCCM} \rightarrow \phi_{\text{O}}^\text{Sn} = 30.1 \text{ nm}^2 \text{s}^{-1}. \\
& \quad \text{(A13)}
\end{align*}
\]

The beam-equivalent pressure (BEP), \( \rho_{\text{r}}^\text{BEP} \), of the reactant fluxes is measured prior to growth by an ion gauge in 10\(^{-7}\) Torr. \( \rho_{\text{r}}^\text{BEP} \) is converted from a pressure into a flux using the kinetic theory of gases [51]. The oxygen flux, \( \phi_{\text{O}}^\text{r} \), is measured in standard cubic centimeters per minute (SCCM) and a radio-frequency plasma power of 300 W is applied. The conversion factors are taken from Ref. [17] and applied to the data given in Ref. [22]. We note that the conversion factors [17] and data [22] are obtained in different MBE systems. Thus, the actual conversion may differ slightly due to the different geometries of the MBE systems used. Nevertheless, the use of the same conversion factors for different MBE systems is a practical and reasonable approach.


P. Mazzolini, A. Falkenstein, C. Wouters, R. Schewski, T. Markurt, Z. Galazka, M. Martin, M. Albrecht, and O. Bierwagen, Substrate-orientation dependence of β-Ga2O3 (100), (010), (001), and (201) homoepitaxy by indium-mediated metal-exchange catalyzed molecular beam epitaxy (MEXCAT-MBE), APL Mater. 8, 011107 (2020).


